

## Acids and Bases



#1

Complete:

$[H^+]$	$[OH^-]$	pH	pOH
		2.4	
			3.8
$1.8 \times 10^{-9}$ mol/L			
	$8.2 \times 10^{-6}$ mol/L		

Solution

$[H^+]$	$[OH^-]$	pH	pOH
$4.0 \times 10^{-3}$ mol/L	$2.5 \times 10^{-12}$ mol/L	2.4	11.6
$6.3 \times 10^{-11}$ mol/L	$1.6 \times 10^{-4}$ mol/L	10.2	3.8
$1.8 \times 10^{-9}$ mol/L	$5.6 \times 10^{-6}$ mol/L	8.7	5.3
$1.2 \times 10^{-9}$ mol/L	$8.2 \times 10^{-6}$ mol/L	8.9	5.1