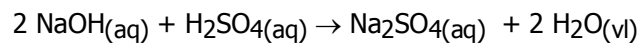


Hoeveel ml  $0,250 \frac{\text{mol}}{\text{L}}$  NaOH-oplossing reageert met 20,0 mL 0,350 M  $\text{H}_2\text{SO}_4$  volgens de reactie



### Oplossing

20,0 mL 0,350 M  $\text{H}_2\text{SO}_4$  bevat  $20,0 \cdot 10^{-3} \text{ L} \cdot 0,350 \frac{\text{mol}}{\text{L}} = 7,00 \cdot 10^{-3} \text{ mol } \text{H}_2\text{SO}_4$

Hiermee reageert  $2 \cdot 7,00 \cdot 10^{-3} \text{ mol} = 1,400 \cdot 10^{-2} \text{ mol } \text{NaOH}$

Deze hoeveelheid is aanwezig in  $\frac{1,400 \cdot 10^{-2} \text{ mol}}{0,250 \frac{\text{mol}}{\text{L}}} = \mathbf{0,056 \text{ L}} = \mathbf{56 \text{ mL}}$   $0,250 \frac{\text{mol}}{\text{L}}$  NaOH-oplossing.