Bereken de molariteit van volgende oplossingen:

- 0,250 mol NaOH in 0,400 L oplossing
- 4,00 g NaOH in 0,400 L oplossing
- 0,250 mol NaOH in 0,800 L oplossing
- 4,00 g NaOH in 1,000 L oplossing

Oplossing

- $c = \dfrac{0,250\ \text{mol}}{0,400\ \text{L}} = \mathbf{0,625}\ \dfrac{\mathbf{mol}}{\mathbf{L}}$

- $c = \dfrac{\dfrac{4,00\ \text{g}}{40,0\ \dfrac{\text{g}}{\text{mol}}}}{0,400\ \text{L}} = \mathbf{0,250}\ \dfrac{\mathbf{mol}}{\mathbf{L}}$

- $c = \dfrac{0,250\ \text{mol}}{0,800\ \text{L}} = \mathbf{0,313}\ \dfrac{\mathbf{mol}}{\mathbf{L}}$

- $c = \dfrac{\dfrac{4,00\ \text{g}}{40,0\ \dfrac{\text{g}}{\text{mol}}}}{1,000\ \text{L}} = \mathbf{0,100}\ \dfrac{\mathbf{mol}}{\mathbf{L}}$